What is the rate chain initiation reaction for initiator efficiency of 0.8, when concentration and decomposition rate constant of benzoyl peroxide are 4 mol/m^3 and $3.24*10^{-6}$ m^3/ mol-s^{-1}, respectively?

The rate expression for chain initiation is given by

$R_i = 2fk_d[I]$
$=2x0.8x3.24*10^{-6}x4$
$=20.736 \times 10^{-6}$ s^{-1}